



# IRON GENIE™

## YOUR IRON & ODOR PROBLEMS SOLVED

- Rust free water
- Iron colored water is virtually gone
- No more staining in stools, tubs, sinks
- Better tasting drinking water

## Iron Genie™ is Simple, Simply the Best!

- No messy or expensive chemicals
- Less maintenance, no filters to change
- Patented product uses no pump or venturi to clog
- Standard Model: TIG56SXT-SI
- Extra Capacity Model: TIG56SXT-SI-ECM

Model Number	Tank Size Width x Height	Media Cu Ft	Max Service* Flow Rate (GPM)	Backwash Rate (GPM)
TIG56SXT-SI	10" x 54"	1	5	7.0
TIG56SXT-SI-ECM	13" x 54"	1.5	8	10.0

\* Max service flow rates are non-continuous for household use only. Flow rates are substantially higher when Zeo Prep filtering media is used.

## How does the Iron Genie™ work?

When water enters the Iron Genie™ it passes through a bubble of compressed air that adds oxygen to it.

The water then passes through a filter bed. The filter material enhances a reaction that separates the iron from the water. The insoluble flakes of iron that result are then caught by the filter bed. (If you'd like to know the chemistry see below!) The iron-free water then flows to your faucets.

The Iron Genie™ replenishes itself every one to three days in a process that washes the iron out from the filter and down the drain, while refilling the compressed air bubble. This happens automatically during the night.

## Why is Iron Genie™ so cost efficient?

- The filter material does not change chemically during the process, so it is not used up and may last indefinitely.
- The other key element in the whole process is the oxygen and that is freely available in the air.
- Finally, the replenishment process is largely driven by the water pressure in your system, so the Iron Genie's™ only running cost is a handful of dollars a year in electricity.

## About the chemistry...

The filter media, which is Manganese Dioxide based, catalyses a reaction that turns the ferrous iron, usually found in water in the form of water-soluble ferrous bicarbonate  $[\text{Fe}(\text{HCO}_3)_2]$ , into ferric oxide ( $\text{Fe}_2\text{O}_3$ ) and ferric hydroxide  $[\text{Fe}(\text{OH})_3]$  which are insoluble. This is the same reaction that takes place when iron-bearing water is exposed to air, which leads to staining of fixtures.

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